

Circular Periodic Table

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ABSTRACT: *The circular form of periodic table is performed on the basis of your increasing positive charged energy level. Here to do something, to remove the some controversial points, consider by International union of pure and applied chemistry. That time the number of elements were limited, so the long form of periodic table is suitable. Now in present it is above to 200. In the long form of periodic table only 118 elements can exist. That time it was also consider that the number of elements would be finite when it is impossible because when time the universe and human body exist new elements will be discovered. The properties of blocks, periods are same as long form of periodic table. There hydrogen exist at only one place but show the position with noble gases as well as metal elements also. The position of lanthanide and actinide series are not showing separate. There exist five block and nine periods if in future any new block and any group will be required they will be exist as define in upper periods as h_1, h_2, \dots etc. So about all type and categories adopted elements can placed in circular periodic table.*

Keywords: { E }, Z, B, G, H, P,

I. INTRODUCTION

Our researcher, scientist, inventors have discovered more than 117 element. 1st by AE beguyer de chancourto is transcribed in 1862 and in 1869 Sir Dmitri Mendleev's manage it according to dependence between the properties of atomic weight, called Mendleev's Periodic table. That time only 63 elements were discovered. In 1892 the long form of periodic table is prepared according to the electronic configuration of element and accepted by IUPAC with some controversial points. In 1915 Seaborg's, extended form (2010). Dirac focu calculation on atoms and ion (Pykko Pekka-2011) were prescribed. In this circular periodic table we to do something to remove some controversial points and make it easy to understand.

II. DESCRIPTION

I study about every field of elements as model and molecular structure, Physical and Chemical properties as Mendleev's, long form, dynamic, extended form, of periodic table etc.

We filtered a common object that the energy level and activity (movement) is main basic point for every element.

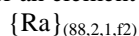
Energy of Element: The energy of element due to consisting group of one type of negative charged energy and positive charged energy with neutron, positron, trans positron etc.

Concept:- Changing is unchanging rule of nature and hope is nature of humanbody. When time the universe (earth, water, air and atmosphere) exist and as well as energy level and weight of elements increased the super heavy, super actinide, super transitive, α -decay, x-ray and other radition, non stable, compactless adopted elements will be discovered in future. The earth is go down (sources) of element, compounds, materials. After some time if any element, compound, material stocks will be finished then that element can be lost. Right now electron, proton, neutron, positron, trans positron, meson, photon, α , β , γ rays, ultraviolet's ray, x-ray, laser rays are discovered, same and different type of above rays will be also discovered in future. The number of elements are not limited. In future many type of elements will be discovered. The S, P, D, F, G, block increased as S, P, D, F, G, H, I, and period will be increased as sp, p₁, p₂, d₁, d₂, f₁, f₂, g₁, g₂, h₁, h₂, etc.

Circular Periodic Table The circular periodic table of element is prepared on the basis of your increasing energy level & according to electronic configuration of element in order to as 2, 8, 8, 18, 18, 32, 32, 50, 50 The digit of circular periodic table distribute every period in equal part or space.

Period:- There consider a period fundamental on the basis of electronic configuration of elements. It is already consider by the scientist that there is different energy level in the form of s, p, d, f, g shell and every shell contains many subshell. The period fundamental is given below-

Show that E is an element. and exist in B-block, G group, H head group and P period contains atomic number Z. Such that consider an element contains atomic number 88, that is Ra (radium) then the figure-



shows that Ra is a element exist in S-block, 2nd group ,first head group & f₂ period contains atomic number 88.

Group :-Theoritically there consider 8 group 1, 2, 3, 4, 5, 6, 7, 8 as a head group, given below-

1. Defined as 1st head group
2. Defined as 2nd head group
3. Defined as 3rd head group
4. Defined as 4th head group
5. Defined as 5th head group
6. Defined as 6th head group
7. Defined as 7th head group
8. Defined as 8th head group

Hydrogen & Helium defined as group superitendant and Hydrogen exist under group 1st, Helium under group 8th.

The group contains elements :-

1st head group in- p₁ period – one element Li atomic number 03.

p₂ period – one element Na atomic number 11.

d₁ period – two elements atomic number 19 to 20.

d₂ period – two elements atomic number 37 to 38.

f₁ period – four elements atomic number 55 to 58.

f₂ period – four elements atomic number 87 to 90.

g₁ period – six elements atomic number 119 to 124.

g₂ period – six elements atomic number 169 to 174.

2nd head group in- p₁ period – one element Be atomic number 4.

p₂ period – one element Ti atomic number 12.

d₁ period – three elements atomic number 21 to 23.

d₂ period – three elements atomic number 39 to 41.

f₁ period – four elements atomic number 59 to 62.

f₂ period – four elements atomic number 91 to 94.

g₁ period – seven elements atomic number 125 to 131.

g₂ period – seven elements atomic number 175 to 181.

3rd head group in- p₁ period – one element B atomic number 5.

p₂ period – one element Be atomic number 13.

d₁ period – two elements atomic number 24 to 25.

d₂ period – two elements atomic number 42 to 43.

f₁ period – four elements atomic number 63 to 66.

f₂ period – four elements atomic number 95 to 98.

g₁ period – seven elements atomic number 132 to 137.

g₂ period – seven elements atomic number 182 to 187.

4th head group in- p₁ period – one element C atomic number 6.

p₂ period – one element Si atomic number 14.

d₁ period – two elements atomic number 26 to 27.

d₂ period – two elements atomic number 44 to 45.

f₁ period – four elements atomic number 67 to 70.

f₂ period – four elements atomic number 99 to 102.

g₁ period – six elements atomic number 138 to 143.

g₂ period – six elements atomic number 188 to 193.

5th head group in- p₁ period – one element N atomic number 7.

p₂ period – one element S atomic number 15.

d₁ period – two elements atomic number 28 to 29.

d₂ period – two elements atomic number 46 to 47.

f₁ period – four elements atomic number 71 to 74.

f₂ period – four elements atomic number 103 to 106.

g₁ period – six elements atomic number 144 to 150.

g₂ period – six elements atomic number 194 to 200.

6th head group in- p₁ period – one element O atomic number 8.

p₂ period – one element P atomic number 16.

d₁ period – three elements atomic number 30 to 32.

- d_2 period – three elements atomic number 48 to 50.
 f_1 period – four elements atomic number 75 to 78.
 f_2 period – four elements atomic number 107 to 110.
 g_1 period – six elements atomic number 151 to 156.
 g_2 period – six elements atomic number 201 to 206.
 7th head group in- p_1 period – one element F atomic number 9.
 p_2 period – one element Cl atomic number 17.
 d_1 period – two elements atomic number 33 to 34.
 d_2 period – two elements atomic number 51 to 52.
 f_1 period – four elements atomic number 79 to 82.
 f_2 period – four elements atomic number 111 to 114.
 g_1 period – six elements atomic number 157 to 162.
 g_2 period – six elements atomic number 207 to 212.
 8th head group in- p_1 period – one element Ne atomic number 10.
 p_2 period – one element Ar atomic number 18.
 d_1 period – two elements atomic number 35 to 36.
 d_2 period – two elements atomic number 53 to 54.
 f_1 period – four elements atomic number 83 to 86.
 f_2 period – four elements atomic number 115 to 118.
 g_1 period – six elements atomic number 163 to 168.
 g_2 period – six elements atomic number 213 to 218.

Miscellaneous group: It is also possible in future that some element will be discovered which have no similarity with any elements, there is a miscellaneous group.

Border Line: The role of border line in the circular periodic table is separate to the group. If any group fill up as your pre-decided limit as show in Fig-1 then the border line advise to go in upper period.

III. CONCLUSION

The circular periodic table is most advantage for chemistry. It is international level solution because about all type & properties adopted elements placed in it and we can already decide the position and nature of new element.

REFERENCE

- [1]. Mendleev's periodic table is prescribed in 1869.
- [2]. Long form of periodic table accepted by IUPAC in 1892.
- [3]. ^{ab}Seaborg (ca. 2006). "transuranium element (chemical element)". Encyclopedia Britannica. Retrieved 2010-03-16.
- [4]. ^ "Extended elements: new periodic table". 2010.
- [5]. ^ ^{ab}Pykkö, Pekka (2011). "A suggested periodic table up to $Z \leq 172$, based on Dirac-Fock calculations on atoms and ions". *Physical Chemistry Chemical Physics* 13 (1): 161-8. Bibcode 2011PCCP...13...161P. doi:10.1039/c0cp01575j. PMID 20967377.
- [6]. ^{ab}Eisberg, R.; Resnick, R. (1985). *Quantum Physics of Atoms, Molecules, Solids, Nuclei and Particles*. Wiley.